

TOPIC – SOLUTIONS

1. Which of the following concentrations is/are independent of temperature?

- a) molality only
- b) molality and molarity
- c) molarity and mole fraction
- d) molality and mole fraction

2. For an aqueous solution the correct statement of the following is

- a) 1M and 1m solutions have same concentration
- b) 1M solution is more concentrated
- c) 1m solution is more concentrated
- d) concentration terms can not be compared

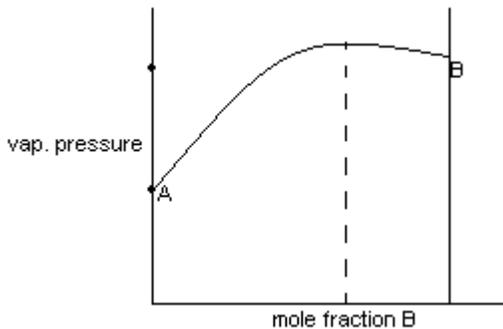
3. When the concentration of salt solution is increased

- a) boiling point increases while vapour pressure decreases
- b) boiling point decreases while vapour pressure increases
- c) freezing point increases while vapour pressure decreases
- d) freezing point decreases while vapour pressure increases

4. Vapour pressure of 1.00 molal NaCl solution is increased by the addition of

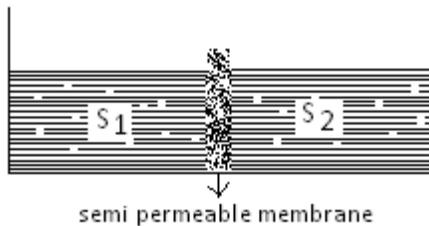
- a) water
- b) KCl
- c) Na₂SO₄
- d) 1.00 molal NaCl

5. The diagram shows vapour pressure - composition for binary solution of A and B. The A - B interactions are



- a) similar to A - A and B - B interactions
- b) greater than A - A and B - B interactions
- c) smaller than A - A and B - B interactions
- d) cannot be predicted

6. Two aqueous solutions S₁ and S₂ are separated as shown in figure.



S₂ has lower vapour pressure than S₁. Then

- a) more solvent flows from S₁ to S₂
- b) more solvent flows from S₂ to S₁

