

Chapter 13 : Concepts in Organic Chemistry

Sl. No.	Question	Obj/ Spec./ Diff. Level
1.	Define electromeric effect.	K Define Easy
	Electromeric effect is a temporary effect in which polarity is induced in a compound with multiple bonds under the requirement of attacking reagent.	
2.	Define inductive effect.	K Define Average
	The permanent displacement of σ electrons along a saturated carbon chain whenever an electron withdrawing or electron releasing group is present at the end of the carbon chain is called as Inductive Effect.	
3.	Mention the reagent used to distinguish between HCHO and C_6H_5CHO .	A Mention Easy
4.	Give an example for the group which shows +M effect.	U Cite example Average
	-Cl or -Br, or -I, or -OH or -NH ₂ or -OCH ₃	
5.	Which among the following groups shows +I effect ? CH ₃ , NO ₂ , COOH, CN	U Identify Average
	- CH ₃	
6.	Which among the following compounds is comparatively stronger base ? a) C ₂ H ₅ NH ₂ b) CH ₃ NH ₂	A Analyse Easy
	C ₂ H ₅ NH ₂	

7. Identify the alkyl group which shows maximum +I effect from the following groups.
- a) $(\text{CH}_3)_3\text{C}-$
 - b) $(\text{CH}_3)_2\text{CH}-$
 - c) CH_3-CH_2-
 - d) $-\text{CH}_3$

8. Distinguish between inductive effect and mesomeric effect.

U
Distinguish
Average

Write any two differences (one mark each)

9. Explain positive inductive effect with an example.

U
Recall
Average

The shared pair of electrons is displaced away from the substituent group (1 mark)

Example (suitable) (1 mark)

10. Give any two differences between inductive and electromeric effect.

U
Distinguish
Average

Write any two differences. (1 + 1 = 2 marks)

11. What is negative inductive effect? Give one example.

K
Recall
Average

The shared pair of σ electrons are displaced towards the substituent.

This is called negative inductive effect. (1 mark)

Ex. Suitable example (1 mark)